1. **Explain, in terms of attractive forces, why there is no increase in temperature of a sample of during melting**
   - Energy is used to break bonds / overcome attraction
   - Between molecules

2. **Why ammonia gas spreads throughout a laboratory**
   - particles move from an area of high to low concentration
   - particles move randomly

3. **Why CO2 diffuses slower than NH3**
   Molecules of CO2 are heavier than molecules of NH3 // CO2 has higher relative molecular mass than NH3

4.

A mixture of ice and water is left to stand and the ice melts.

Which row describes what happens as the ice is melting?

	temperature of mixture	energy changes
A	increases	average kinetic energy of particles increases
B	increases	energy is used to overcome attractive forces
C	stays the same	average kinetic energy of particles increases
D	stays the same	energy is used to overcome attractive forces

Answer: D

5. **How motion of particles of a solid changes when it is heated (before reaching m.p)**
   - Vibrations
   - Increase

6. **Why do gases diffuse**
   Due to random motion of molecules/ particles